

4541/2
Chemistry
Kertas 2
2012
2 $\frac{1}{2}$ jam

NAMA :

NO. KAD PENGENALAN :

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**PEPERIKSAAN PERCUBAAN BERSAMA
SIJIL PELAJARAN MALAYSIA 2012**

**ANJURAN
MAJLIS PENGETUA SEKOLAH-SEKOLAH MALAYSIA
(MPSM) CAWANGAN NEGERI PERLIS**

CHEMISTRY

Kertas 2

Dua jam tiga puluh minit

JANGAN BUKA KERTAS SOALAN INI SEHINGGA DIBERITAHU

1. Tuliskan nombor kad pengenalan dan nama anda pada ruang yang disediakan.
2. Kertas soalan ini adalah dalam dwibahasa.
3. Soalan dalam bahasa Inggeris mendahului soalan yang sepadan dalam bahasa Melayu.
4. Calon dibenarkan menjawab keseluruhan atau sebahagian soalan sama ada dalam bahasa Melayu atau bahasa Inggeris.
5. Calon dikehendaki membaca maklumat di halaman 25

Kod Pemeriksa			
Bahagian	Soalan	Markah Penuh	Markah Diperoleh
A	1	9	
	2	9	
	3	10	
	4	10	
	5	11	
	6	11	
B	7	20	
	8	20	
C	9	20	
	10	20	
Jumlah			

Kertas soalan ini mengandungi 28 halaman bercetak

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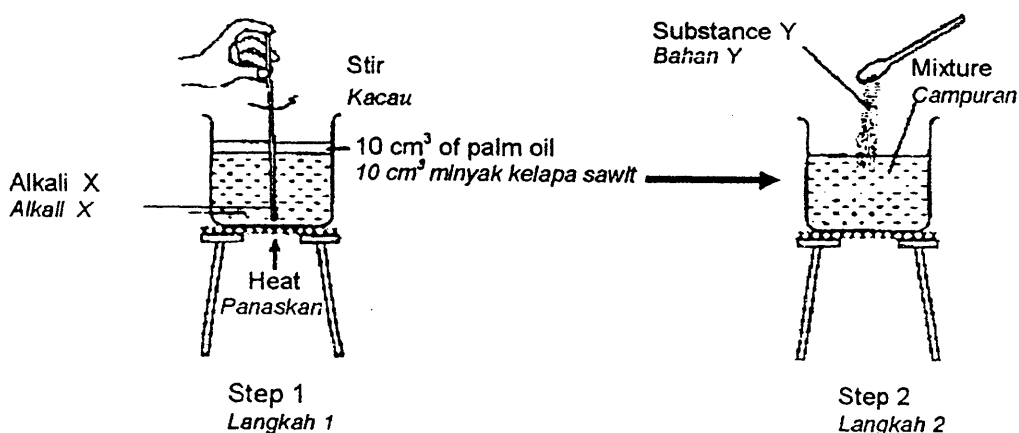
Section A
Bahagian A

[60 Marks]

[60 Markah]

Answer all questions in this section.
Jawab semua soalan dalam bahagian ini.

- 1 Diagram 1 shows the apparatus set-up for preparing soap.
Rajah 1 menunjukkan susunan radas untuk menyediakan sabun.



- (a) Name the process.
Namakan proses itu.

[1 mark]
[1 markah]

- (b) A student wants to prepare sodium palmitate soap.
Name alkali X used in step 1
Seorang pelajar ingin menyediakan sabun kalium palmitat.
Namakan alkali X yang digunakan dalam langkah 1.

[1 mark]
[1 markah]

- (c) (i) Substance Y is added in the mixture in step 2.
Name substance Y
Bahan Y dimasukkan kedalam campuran dalam langkah 2
Namakan bahan Y.

[1 mark]
[1 markah]

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- (ii) State why substance Y is added to the mixture.
Nyatakan mengapa bahan Y ditambah kepada campuran itu.

[1 mark]

[1 markah]

- (d) Diagram 1.2 shows the structure of the anion of a soap molecule.
Rajah 1.2 menunjukkan struktur anion molekul sabun.

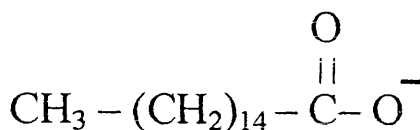


Diagram 1.2
Rajah 1.2

- (i) Circle the hydrophilic part, in Diagram 1.2
Bulatkan bahagian hidrofilik, dalam Rajah 1.2.

[1 mark]

[1 markah]

- (ii) Which part is soluble in grease?
Bahagian manakah boleh larut dalam gris?

[1 mark]

[1 markah]

- (iii) Clothes becomes clean and does not form scum when wash with soap. Mark (✓) in the box in Table 1.3 to the correct type of water used.
Pakaian menjadi bersih dan tidak membentuk kekat apabila dicuci dengan sabun. Tandakan (✓) dalam kotak dalam Jadual 1.3 untuk menunjukkan jenis air yang digunakan.

Hard water <i>Air liat</i>	Soft water <i>Air lembut</i>

Table 1.3
Jadual 1.3

[1 mark]

[1 markah]

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- (e) Table 1.4 shows the functions of two types of food additives
Jadual 1.4 menunjukkan fungsi bagi dua jenis bahan tambah makanan.

Function <i>Fungsi</i>	Type of Food Additives <i>Jenis Bahan Tambah Makanan</i>
To prevent growth of microorganism. <i>Menghalang pertumbuhan mikroorganisma.</i>	P:
To prevent oxidation that causes rancid fats and brown food. <i>Menghalang pengoksidaan yang menyebabkan lemak tengik dan makanan berwarna perang.</i>	Q:

Table 1.4
Jadual 1.4

Complete the Table 1.4
Lengkapkan Jadual 1.4

[2 marks]
[2 markah]

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2.

- (a) Diagram 2 shows the standard representation of two isotopes of chlorine atoms.
Rajah 2 menunjukkan perwakilan atom bagi dua isotop atom klorin.

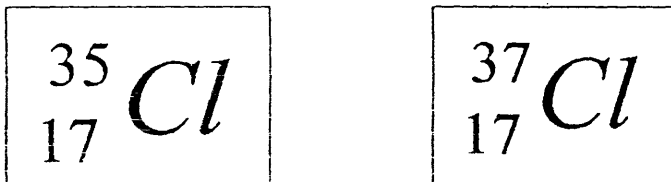


Diagram 2
Rajah 2

- (i) Based on Diagram 2, state the meaning of isotope.
Berdasarkan Rajah 2, nyatakan maksud isotop.

.....
.....

[2 marks]
[2 markah]

- (ii) Determine the number of neutrons in
Tentukan bilangan neutron dalam

$\begin{matrix} 35 \\ 17 \end{matrix} Cl$:

$\begin{matrix} 37 \\ 17 \end{matrix} Cl$:

[2 marks]
[2 markah]

- (iii) State one of the uses of chlorine in our daily lives.
Nyatakan satu kegunaan klorin dalam kehidupan seharian.

.....
[1 mark]
[1 markah]

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- (b) X is a substance that has melting point of 43°C and boiling point of 89°C .
X adalah satu bahan yang mempunyai takat lebur 43°C dan takat didih 89°C .
- (i) Sketch a graph of temperature against time when substance X is heated from 30°C to 80°C .
Lakarkan graf suhu melawan masa apabila bahan X dipanaskan daripada 30°C hingga 80°C .

[2 marks]
[2 markah]

- (ii) Explain the arrangement of particles in substance X at :
Terangkan susunan zarah dalam bahan X pada :

30°C :

.....
.....

80°C :

.....
.....

[2 marks]
[2 markah]

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- 3 Nitric acid is a strong acid and ethanoic acid is a weak acid. Table 3 shows the concentration of each acid solution.
Asid nitrik adalah asid kuat dan asid etanoik adalah asid lemah. Jadual 3 menunjukkan kepekatan setiap larutan asid..

Acid Asid	Concentration (mol dm^{-3}) Kepekatan (mol dm^{-3})
Nitric acid solution Larutan asid nitrik	0.10
Ethanoic acid solution Larutan asid etanoik	0.10

Table 3
Jadual 3

- (a) State the meaning of acid.
Nyatakan maksud asid.

[1 mark]
[1 markah]

- (b) Give a name of another strong acid.
Berikan satu nama asid kuat yang lain.

[1 mark]
[1 markah]

- (c) Nitric acid solution and ethanoic acid solution in Table 3 have different pH values.
Larutan asid nitrik dan larutan asid etanoik dalam Jadual 3 mempunyai nilai pH yang berbeza.

- (i) Which solution gives a higher pH value?
Larutan yang manakah memberi nilai pH yang lebih tinggi?

[1 mark]
[1 markah]

- (ii) Give one reason for the answer in 3(c)(i).
Beri satu sebab bagi jawapan di 3(c)(i).

[1 mark]
[1 markah]

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- (d) 30 cm^3 of 0.1 mol dm^{-3} potassium hydroxide solution is poured in a conical flask. Then a few drops of phenolphthalein indicator solution are added. The solution is titrated with nitric acid solution in Table 3.

30 cm^3 larutan kalium hidroksida 0.1 mol dm^{-3} dimasukkan ke dalam kelalang kon. Kemudian beberapa titis larutan penunjuk fenolftalein ditambah. Larutan ini dititratkan dengan larutan asid nitrik dalam Jadual 3.

- (i) State the type of reaction between potassium hydroxide solution and nitric acid solution.
Nyatakan jenis tindak balas antara larutan kalium hidroksida dengan larutan asid nitrik.

.....
[1 mark]
[1 markah]

- (ii) What is the colour change of the mixture at the end point?
Apakah perubahan warna campuran itu pada takat akhir?

.....
[1 mark]
[1 markah]

- (iii) Write the chemical equation for the reaction.
Tulis persamaan kimia bagi tindak balas itu.

.....
[2 marks]
[2 markah]

- (iv) Calculate the volume of the nitric acid used.
Hitung isipadu asid hidroklorik yang digunakan.

.....
[2 marks]
[2 markah]

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- 4 A student carried out two experiments to investigate the effect of concentration of hydrochloric on the rate of reaction.. Table 4 shows the results of the experiments.

Seorang pelajar telah menjalankan dua eksperimen untuk menyiasat kesan kepekatan asid hidroklorik ke atas kadar sesuatu tindak balas. Jadual 4 menunjukkan keputusan eksperimen.

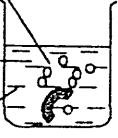
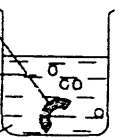
Experiment	Diagram	Time taken for all magnesium dissolve/s Masa yang diambil untuk semua magnesium melarut/s
I	<p>1.2 g magnesium ribbon 1.2 g pita magnesium</p> <p>Gas bubbles Gelembung gas</p>  <p>Excess hydrochloric acid 1 mol dm⁻³ asid hidroklorik 1 mol dm⁻³ berlebihan</p>	30
II	<p>1.2 g magnesium ribbon 1.2 g pita magnesium</p> <p>Gas bubbles Gelembung gas</p>  <p>Excess hydrochloric acid 0.5 mol dm⁻³ asid hidroklorik 0.5 mol dm⁻³ berlebihan</p>	60

Table 4
Jadual 4

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- (a) (i) Name the gas released in the experiment
Namakan gas yang terbebas dalam eksperimen ini
-
- [1 mark]
[1 markah]
- (ii) Write a chemical equation for the reaction between magnesium and hydrochloric acid.
Tulis persamaan kimia bagi tindak balas antara magnesium dengan asid hidroklorik.
-
- [2 marks]
[2 markah]
- (iii) Calculate the maximum volume of the gas released at room temperature in this experiment.
[Relative atomic mass : Mg,24;
Volume of 1 mole of gas at room temperature = 24 dm³]
*Hitung isipadu maksimum gas yang terbebas pada suhu bilik dalam eksperimen ini.
[Jisim atom relatif : Mg,24;
Isipadu 1mol gas pada suhu bilik = 24 dm³]*
-
- [3 marks]
[3 markah]
- (b) Compare the rate of reaction between experiment I and experiment II. Explain.
Bandingkan kadar tindak balas antara eksperimen I dan eksperimen II. Terangkan
-
-
- [2 marks]
[2 markah]
- (c) State two other factors that can affect the rate of reaction in this eksperimen.
Nyatakan dua faktor lain yang boleh mempengaruhi kadar tindak balas dalam eksperimen ini.
-
-
- [2 marks]
[2 markah]

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5. Diagram 5 shows two types of electrochemical cells.
Rajah 5 menunjukkan dua jenis sel elektrokimia.

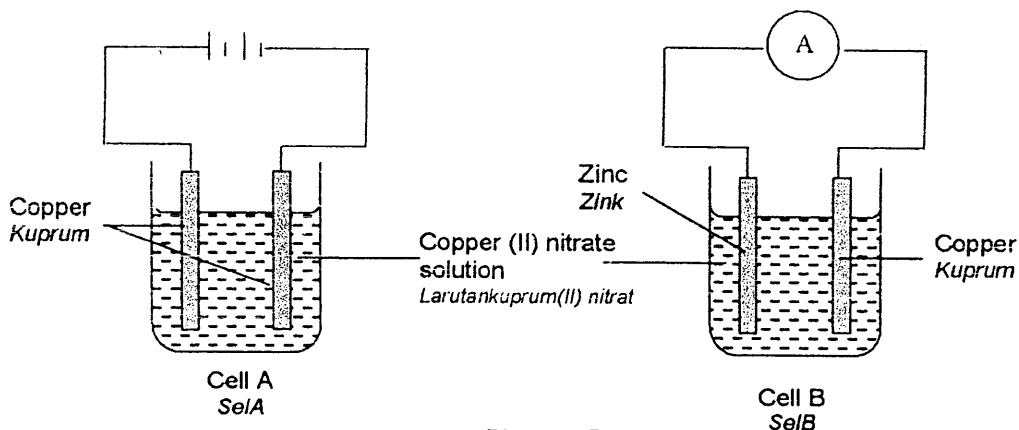


Diagram 5
Rajah 5

- (a) State all the anion present in the copper(II) nitrate solution.
Nyatakan semua anion yang terdapat dalam larutan kuprum(II) nitrat.

[1 mark]
[1 markah]

- (b) (i) Name the type of Cell A.
Namakan jenis Sel A.

[1 mark]
[1 markah]

- (ii) State one use of Cell A in industry.
Nyatakan satu kegunaan Sel A di dalam industri.

[1 mark]
[1 markah]

- (c) (i) State the observation at the anode in Cell A.
Nyatakan pemerhatian pada anod dalam Sel A.

[1 mark]
[1 markah]

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- (ii) Based on the answer at (c)(i) write the half equation for the reaction occurs at anode in Cell A.
Berdasarkan jawapan pada(c)(i) tuliskan setengah persamaan bagi tindak balas yang berlaku di anode dalam Sel A.

.....
[1 mark]
[1 markah]

- (d) Name the product formed at the cathode if copper electrodes in Cell A are replaced by carbon electrodes.
Namakan hasil yang terbentuk di katod jika elektrod kuprum dalam Sel A diganti oleh elektrod karbon.

.....
[1 mark]
[1 markah]

- (e) Based on Cell B,
Berdasarkan Sel B,

- (i) show the flow of electrons direction in Cell B.
tunjukkan arah pergerakan elektron dalam Sel B.

[1 mark]
[1 markah]

- (ii) State the negative terminal and positive terminal.
Nyatakan terminal negatif dan terminal positif.

Negative terminal:
Terminal negatif :

.....
Positive terminal:
Terminal positif :

.....
[2 marks]
[2 markah]

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- (f) State the colour change of copper(II) nitrate solution in Cell B.
Explain your answer.
*Nyatakan perubahan warna larutan kuprum(II) nitrate dalam Sel B.
Terangkan jawapan anda.*

.....

.....

.....

.....

[2 marks]
[2 markah]

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6. The following information is about compound C_2H_4 .
Maklumat berikut adalah mengenai sebatian C_2H_4 .

- | |
|--|
| <ul style="list-style-type: none">▪ Gas at room temperature
<i>Gas pada suhu bilik</i>▪ Undergoes combustion
<i>Menjalani pembakaran</i>▪ Member of a homologous series
<i>Ahli suatu siri homolog</i> |
|--|

- (a) What is the name of this compound?
Apakah nama bagi sebatian ini?

.....
[1 mark]
[1 mark]

- (b) State the general formula and the functional group for the homologous series of this compound.
Nyatakan formula am dan kumpulan berfungsi bagi siri homologous sebatian ini?

.....
[2 marks]
[2 mark]

- (c) Compound C_2H_4 burns in excess oxygen. Write the chemical equation for this reaction. *Sebatian C_2H_4 terbakar dalam oksigen berlebihan. Tuliskan persamaan kimia bagi tindak balas ini.*

.....
[2 marks]
[2 mark]

- (d) Ethanol undergoes dehydration to produce compound C_2H_4 .
Etanol mengalami pendehidratan untuk menghasilkan sebatian C_2H_4 .

- (i) Name one dehydrating agent for this reaction.
Namakan satu agen pendehidartan bagi tindak balas ini.

.....
[1 mark]
[1 mark]

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- (ii) Draw the apparatus set-up diagram for the dehydration reaction that can be used to collect compound C_2H_4 .
Lukiskan gambarajah susunan radas untuk tindak balas penghidratan yang boleh digunakan untuk mengumpul sebatian C_2H_4

[2 marks]
 [2 mark]

- (e) Table 6 shows the results of a test to differentiate between compound C_2H_4 and ethane.
Jadual 6 menunjukkan keputusan ujian untuk membezakan antara sebatian C_2H_4 dan etana.

Procedure <i>Prosedur</i>	Observations <i>Pemerhatian</i>
Compound C_2H_4 flow into bromine water <i>Sebatian C_2H_4 dialirkan kedalam air bromin</i>	Brown gas is decolourised <i>Warna perang dinyahwarnakan</i>
Ethane flow into bromine water <i>Etana dialirkan ke dalam air bromin</i>	Brown colour remains <i>Warna perang tidak berubah</i>

Table 6
 Jadual 6

- (i) Explain why there is a difference in this observations.
Terangkan mengapa terdapat perbezaan dalam pemerhatian

[1 mark]
 [1 mark]

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- (ii) Draw the structural formula of the compound formed when compound C_2H_4 is added into bromine water.
Lukis formula struktur bagi sebatian yang terbentuk apabila sebatian C_2H_4 dialirkan ke dalam air bromin

[1 mark]
[1 mark]

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Section B
 Bahagian B
 [20 marks]
 [20 markah]

Answer any one question.
 Jawab mana-mana satu soalan.

7

- (a) Table 7.1 shows the electron arrangements of atoms W, X and Y. These letters are not the actual symbols of the elements.
 Jadual 7.1 menunjukkan susunan elektron bagi atom W, X dan Y. Huruf-huruf ini bukanlah simbol sebenar bagi unsur-unsur tersebut.

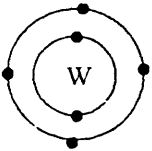
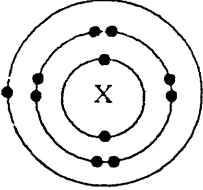
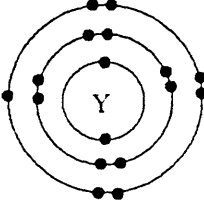
Element unsur	Electron arrangement Susunan elektron
W	
X	
Y	

Table 7.1
 Jadual 7.1

Based on the table,

- (i) State one element that is metal and one element that is non-metal.
 Nyatakan satu unsur logam dan satu unsur bukan logam.
 [2 marks]
 [2 markah]
- (ii) State the position of element W in the Periodic table of Element. Explain your answer.
 Nyatakan kedudukan unsur W dalam Jadual Berkala Unsur. Jelaskan jawapan anda.
 [3 marks]
 [3 markah]

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- (ii) Compare the atomic size of element X and element Y. Explain your answer.
Bandingkan saiz atom unsur X dan unsur Y. Terangkan jawapan anda.

[4 marks]

[4 markah]

- (b) A, D and E are placed in Period 3 of the Periodic Table of Elements. The oxides of these elements show different properties when react with nitric acid and sodium hydroxide solution. Table 7.2 shows the properties of the oxides of the elements.
Unsur-unsur A, D dan E berada dalam Kala 3 Jadual Berkala Unsur. Oksida-oksida unsur tersebut menunjukkan sifat yang berbeza apabila bertindak balas dengan asid nitrik dan larutan natrium hidroksida. Jadual 7.2 menunjukkan sifat-sifat oksida bagi unsur-unsur tersebut.

Oxide Oksida	Observation Pemerhatian	
	Reaction with nitric acid <i>Tindak balas dengan asid nitrik</i>	Reaction with sodium hydroxide solution <i>Tindak balas dengan larutan natrium hidroksida</i>
Oxide of Element A Oksida unsur A	White solid dissolves to form colourless solution <i>Pepejal putih larut dan membentuk larutan tidak berwarna</i>	No change. White solid does not dissolve. <i>Tiada perubahan. Pepejal putih tidak larut</i>
Oxide of Element D Oksida unsur D	White solid dissolves to form colourless solution. <i>Pepejal putih larut dan membentuk larutan tidak berwarna</i>	White solid dissolves to form colourless solution <i>Pepejal putih larut dan membentuk larutan tidak berwarna</i>
Oxide of Element E Oksida unsur E	No change. White solid does not dissolve. <i>Tiada perubahan. Pepejal putih tidak larut</i>	White solid dissolves to form colourless solution. <i>Pepejal putih larut dan membentuk larutan tidak berwarna</i>

Table 7.2
 Jadual 7.2

Based on table 7.2,
 Berdasarkan Jadual 7.2,

- (i) State the acid-base properties of the oxide of A, oxide of D and oxide of E.
Nyatakan sifat-sifat asid-bes bagi oksida A, oksida D dan oksida E. Terangkan.

[6 marks]

[6 markah]

- (ii) Suggest the name of element A, element D and element E.
Cadangkan nama bagi unsur A, unsur D dan unsur E.

[3 marks]

[3 markah]

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- (iii) Write the chemical equation for the reaction between oxide of A and nitric acid

Tuliskan persamaan kimia bagi tindak balas antara oksida A dan asid nitrik.

[2 marks]

[2 markah]

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8. Table 8 shows thermochemical equations of Experiment 1 and Experiment 2.
Jadual 8 menunjukkan persamaan termokimia bagi Eksperimen 1 dan Eksperimen 2.

Experiment Eksperimen	Thermochemical Equation Persamaan Termokimia
1	$\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O} \quad \Delta H = -57.3 \text{ kJ mol}^{-1}$
2	$\text{NaOH} + \text{CH}_3\text{COOH} \rightarrow \text{CH}_3\text{COONa} + \text{H}_2\text{O} \quad \Delta H = -51.5 \text{ kJ mol}^{-1}$

Table 8
Jadual 8

- (a) Based on Table 8,
Berdasarkan Jadual 8,

- (i) state the type of reaction.
nyatakan jenis tindak balas

[1 mark]
[1 markah]

- (ii) state the differences between the total energy content of the reactants and the total energy content of products.
nyatakan perbezaan di antara kandungan tenaga bahan tindak balas dan kandungan tenaga hasil tindak balas.

[1 mark]
[1 markah]

- (iii) Explain the difference of the heat of neutralization for the reactions above.
Jelaskan perbezaan haba peneutralan bagi tindak balas–tindak balas di atas.

[5 marks]
[5 markah]

- (b) In Experiment 1, 50 cm^3 of 1.0 mol dm^{-3} sodium hydroxide solution is reacted with 50 cm^3 of 1.0 mol dm^{-3} of hydrochloric acid solution.
Di dalam Eksperimen 1, 50 cm^3 larutan natrium hidroksida, 1.0 mol dm^{-3} ditindak balaskan dengan 50 cm^3 larutan asid hidroklorik, 1.0 mol dm^{-3} .

- (i) Calculate the heat released in the reaction.
Kirakan haba yang terbebas dari tindak balas.

[3 marks]
[3 markah]

- (ii) What is the temperature change of the reaction mixture?
[Specific heat capacity = $4.2 \text{ J g}^{-1} \text{ } ^\circ\text{C}^{-1}$, density of solution = 1 g cm^{-3}]
Berapakah perubahan suhu campuran tindak balas?
[Muatan haba tentu larutan = $4.2 \text{ J g}^{-1} \text{ } ^\circ\text{C}^{-1}$, ketumpatan larutan = 1 g cm^{-3}]

[3 marks]
[3 markah]

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- (iii) The heat of neutralisation obtained in laboratory is less than the actual heat of neutralisation. Explain.
Nilai haba peneutralan yang diperolehi di makmal adalah lebih rendah dari nilai haba peneutralan yang sebenar. Jelaskan.

[2 marks]
[2 markah]

- (c) Diagram 8.2 shows two energy level diagrams of different reactions.
Rajah 8.2 menunjukkan dua gambarajah aras tenaga bagi tindak balas berlainan.

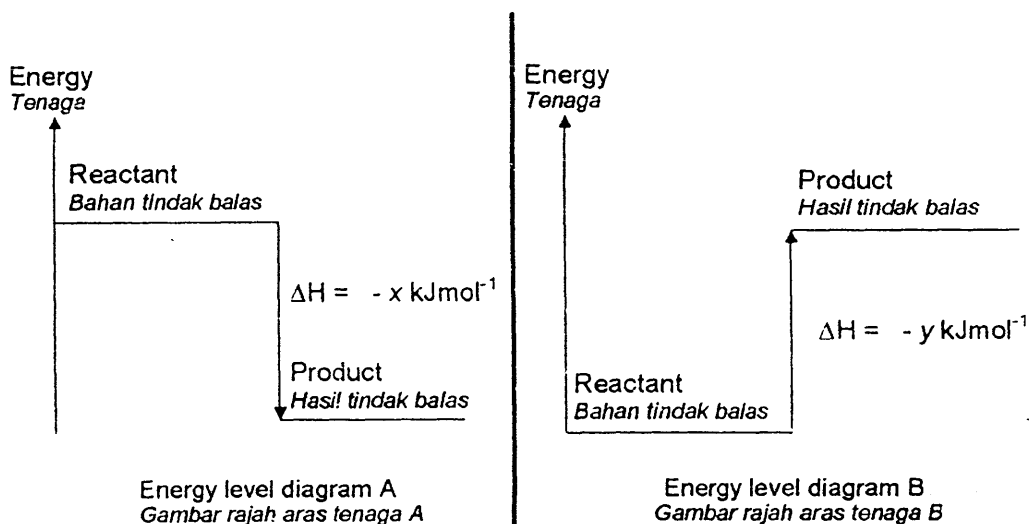


Diagram 8.2
Rajah 8.2

Based on Diagram 8.2, compare the energy level diagram A and energy level diagram B.
Berdasarkan Rajah 8.2, bandingkan gambarajah aras tenaga A dan gambarajah aras tenaga B.

[5 marks]
[5 markah]

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Section C
Bahagian C
[20 marks]
[20 markah]

Answer any one question.
Jawab mana-mana *satu* soalan.

9

- (a) Diagram 9 shows one of the stages involved in the Contact Process.
Rajah 9 menunjukkan salah satu peringkat dalam Proses Sentuh.

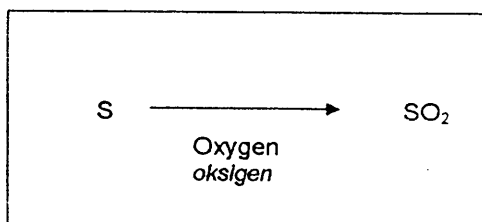


Diagram 9
Rajah 9

- (i) Write the chemical equation for the reaction.
Tuliskan persamaan kimia bagi tindakbalas tersebut.

[2 marks]
[2 markah]

(ii)

Sulphur dioxide is the waste product from the Contact Process which affects the quality of the environment.
Sulfur dioksida adalah bahan buangan daripada Proses Sentuh yang menjejaskan kualiti alam sekitar.

As a chemist, how do you advice the manager of the factory to reduce the release of sulphur dioxide gas. Explain.

Sebagai seorang ahli kimia, bagaimana anda menasihati pengurus kilang tersebut supaya pembebasan gas sulfur dioksida dikurangkan. Terangkan.

[2 marks]
[2 markah]

SULIT

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- (b) Table 9 shows three different manufactured substances in industry X, Y, and Z, and their uses.
Jadual 9 menunjukkan tiga bahan buatan Industri X, Y, dan Z, yang berbeza dan kegunaannya.

Manufactured substances in industry <i>Bahan buatan Industri</i>	Uses <i>Kegunaan</i>
X	To make internal wall of furnace <i>Untuk membuat lapisan dalam dinding relau</i>
Y	To make musical instruments <i>Untuk membuat alat muzik</i>
Z	To make helmets and water storage tanks <i>Untuk membuat topi keledar dan tangki penyimpanan air</i>

Table 9
Jadual 9

Based on Table 9, state the name of X, Y and Z.
Give the specific properties of each of the substances to support your answers.
Berdasarkan Jadual 9, nyatakan nama bagi X, Y dan Z.
Berikan sifat khusus bagi setiap bahan tersebut untuk menyokong jawapan anda.

[6 marks]
[6 markah]

- (c) The following information shows two properties of alloys.
Maklumat berikut menunjukkan dua sifat aloi.

- Alloy is harder than its pure metal
Aloi lebih keras daripada logam tulennya
- Alloy is more corrosive resistant than its pure metal
Aloi lebih tahan kakisan daripada logam tulennya

By choosing one of the properties of the alloy stated above, describe a laboratory experiment to compare alloy with its pure metal.

Your answer should include the following:

- Name of the alloy and its pure metal
- Materials and apparatus
- Procedure of the experiment
- Observation
- Conclusion

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Dengan memilih satu daripada sifat aloi yang dinyatakan diatas, huraikan satu eksperimen makmal untuk membandingkan aloi dengan logam tulennya.

Jawapan anda perlu mengandungi perkara berikut:

- *Nama aloi dan logam tulennya*
- *Bahan dan radas*
- *Kaedah eksperimen*
- *Pemerhatian*
- *kesimpulan*

{10 marks}
{10 markah}

SULIT

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10

- (a) 50 cm^3 of 0.2 mol dm^{-3} sulphuric acid reacts with excess copper(II) oxide to produce copper(II) sulphate and water..
 50 cm^3 asid sulfurik 0.2 mol dm^{-3} bertindak balas dengan kuprum(II) oksida yang berlebihan menghasilkan kuprum(II) sulfat dan air.

- (i) Write the chemical equation for the reaction that takes place.
Tuliskan persamaan kimia bagi tindak balas yang berlaku.
- (ii) Calculate the mass of copper(II) sulphate that is produced
Hitungkan jisim kuprum(II) sulfat yang terhasil.
[Molar mass of copper(II) sulphate = 160 g mol^{-1}]
[Jisim molar kuprum(II) sulfat = 160 g mol^{-1}]

[4 marks]
[4 markah]

- (b) The following are two examples of chloride salts that can be prepared in the laboratory.
Berikut adalah dua contoh garam klorida yang boleh disediakan di dalam makmal.

- | |
|--|
| <ul style="list-style-type: none">▪ Potassium Chloride, KCl
<i>Kalium klorida, KCl</i>▪ Silver chloride, AgCl
<i>Argentum klorida, AgCl</i> |
|--|

Compare and contrast between these two salts. Your comparison should include the following aspects:

- the solubility of each salt in water.
- the name of the chemical reaction for the preparation of each salt.
- the reactants for the preparation of each salt.
- the chemical equation for preparation of each salt.

Banding dan bezakan diantara dua garam ini. Perbandingan anda hendaklah mengandungi aspek-aspek berikut:

- keterlarutan setiap garam di dalam air.
- nama tindak balas kimia untuk penyediaan setiap garam.
- bahan tindak balas untuk penyediaan setiap garam.
- persamaan kimia untuk penyediaan setiap garam.

[6 marks]
[6 markah]

SULIT

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- (c) Barium sulphate is insoluble in water. Describe the preparation of barium sulphate in the laboratory. In your description, include

- reactants
- procedure
- chemical equations involved

Barium sulfat adalah tak larut dalam air. Huraikan penyediaan barium sulfat dalam makmal. Dalam huraian anda, sertakan

- *bahan tindak balas*
- *prosedur*
- *persamaan kimia yang terlibat*

[10 marks]
[10 markah]

**END OF QUESTION PAPER
KERTAS SOALAN TAMAT**

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Perlis

2

SPM Kimia Paper (2) Jawapan 2012

1.	(a)	(i)	Saponification	1				
	(b)		Sodium hydroxide	1				
	(c)	(i)	Sodium chloride	1				
		(ii)	To precipitate the soap/to reduce solubility of soap	1				
	(d)	(i)	<div>$\text{CH}_3 - (\text{CH}_2)_{14} - \text{C} \begin{array}{c} \text{O} \\ \\ \text{O}^- \end{array}$</div>	1				
	(ii)	Hydrophobic part	1					
	(iii)	<table><tr><td>Hard water <i>Air liat</i></td><td>Soft water <i>Air lembut</i></td></tr><tr><td></td><td>✓</td></tr></table>		Hard water <i>Air liat</i>	Soft water <i>Air lembut</i>		✓	1
Hard water <i>Air liat</i>	Soft water <i>Air lembut</i>							
	✓							
(e)	P : preservatives			1				
	Q : antioxidants			1				
	TOTAL			9				

2	(a)	i)	Same element that have same proton number but different nucleon number	1
		ii)	Cl-35:18	1
			Cl-37:20	1
		iii)	Used in water treatment to kill microorganisms/germs	1
	(b)	i)	<p>Temperature (°C)</p> <p>Time(s)</p>	1
		ii)	<p>At 30 °C : the particles are arranged closely packed in orderly manner</p> <p>At 80 °C : the particles are arranged closely packed but not in orderly manner</p>	1
	TOTAL			9

3	(a)	The chemical substance that ionizes in water to produce hydrogen ion / H^+ .	1	
		(b) Nitric acid	1	
	(b)	i) Hydrochloric acid	1	
		ii) Hydrochloric acid produce lower concentration of hydrogen ion / H^+ .	1	
		iii) Neutralisation reaction	1	
	(c)	i) Pink to colourless	1	
		ii) $HNO_3 + KOH \longrightarrow KNO_3 + H_2O$	1+1	
		$0.1 \times V_A / 0.1 \times 30 = 1 / 1$ $V_A = 30 \text{ cm}^3$	1 1	
			Total	10

4	(a)	(i)Hydrogen	1
		(ii) $Mg + 2HCl \longrightarrow MgCl_2 + H_2$	1
		- Reactants and products are correct	1
		- Equation is balanced	
	(iii) No of mol Mg = $1.2/24 = 0.05 \text{ mol}$ Mol of $H_2 = 0.05 \text{ mol}$ Volume of $H_2 = 0.05 \times 24 \text{ dm}^3/1.2 \text{ dm}^3/1200 \text{ cm}^3$	1 1 1	
(b)	- Experiment I is higher - Concentration of hydrochloric acid is higher/Time taken in experiment I is shorter ** Reject if not mentioned EXP I	1 1	
(c)	-Uses of catalyst - Temperature	1 1	
	TOTAL		10

5	(a)	NO_3^- , OH^-	1
	(b)	(i) Electrolytic cell	1
		(ii) Electroplating/Purification	1
	(c)	(i) Thinner	1
		(ii) $Cu^{2+} + 2e \rightarrow Cu$	1
	(d)	Copper//Copper atom/metal	1
	(e)	(i) Students label the arrow of electron flow in the diagram (From zinc plate to copper plate terminal)	1
		(ii) Negative terminal: Zinc	1

Perlis

4

SPM Kimia Paper (2) Jawapan 2012

		Positive terminal: Copper	1
(f)		Blue solution change to colourless//blue solution becomes paler//blue solution decolourised	1
		Because the concentration of Cu^{2+} ion in the solution decreases	1
		TOTAL	11

6	(a)	Ethene	1
	(b)	<p>- Alkene</p> <p>-Carbon- carbon double bond/</p> $\begin{array}{c} \quad \\ -\text{C}=\text{C}- \\ \quad \end{array}$	1 1
	(c)	<p>(i) $\text{C}_2\text{H}_4 + 3\text{H}_2\text{O} \longrightarrow 3\text{CO}_2 + 2\text{H}_2\text{O}$</p> <p>- correct reactants and products</p> <p>- balanced equation</p>	1 1
	(d)	<p>(i) Porous pot/Concentrated sulfuric acid /Porcelain chips/Aluminium oxide</p> <p>(ii)</p> <div style="border: 1px solid black; padding: 10px; margin: 10px;"> </div> <p>-Functional diagram</p> <p>-Labelled diagram</p>	1 1 1 1
	(e)	<p>(i) - C_2H_4 is an unsaturated hydrocarbon/ has double bond</p> <p>- Ethane is a saturated hydrocarbon/ has single bond</p> <p>(ii)</p> $\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{H} \\ \quad \\ \text{Br} \quad \text{Br} \end{array}$	1 1 1
		TOTAL	11

Section B

SECTION B

7	(a)	(i) metal: X non-metal: W//Y	1 1
		(ii) Electron arrangement of W: 2.4 Period 2 because atom W has 2 shells filled/occupied with electrons Group 14 because atom W has 4 valence electron	1 1 1
		(iii) <ul style="list-style-type: none"> Atomic size of atom Y is smaller The positive charge /the number of proton in the nucleus of atom Y is bigger Force of attraction between the nucleus and the electrons in atom Y is stronger The electrons are pulled closer to the nucleus 	1 1 1 1
	(b)	(i) Oxide of A: <u>Basic</u> because <u>oxide of A react with acid only</u> to form salt and water Oxide of D: <u>Amphoteric</u> because <u>oxide of D react with both acid and base</u> to form salt and water. Oxide of E: <u>Acidic</u> because <u>oxide of E react with base only</u> to form salt and water. (ii) A: Sodium/Magnesium D: Aluminium E: Silicon/Sulphur/Chlorine/Phosphorus (iii) $\text{Na}_2\text{O} + 2\text{HNO}_3 \longrightarrow 2\text{NaNO}_3 + \text{H}_2\text{O}$ // $\text{MgO} + 2\text{HNO}_3 \longrightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2\text{O}$	1 + 1 1 + 1 1 + 1 1 1 1 1 + 1
		Total	20

Perlis

6

SPM Kimia Paper (2) Jawapan 2012

8	(a)	Neutralisation//Exothermic reaction	1												
	(i)														
	(ii)	Total energy content of reactant is higher than total energy content in product	1												
	(iii)	<ul style="list-style-type: none">The heat of neutralization of Expt 1 is higher than Expt 2HCl is strong acid while ethanoic acid is weak acidHCl ionizes completely in water to produce H⁺ ionCH₃COOH ionizes partially in water to produce H⁺ ion and most of ethanoic acid exist as moleculesIn Expt 2, Some of heat given out during neutralization reaction is used to dissociate the ethanoic acid completely in water//part of heat that is released is used to break the bonds in the molecules of ethanoic acid that has not been ionised	1 1 1 1 1												
	(b)														
	(i)	No of mol acid/alkali= 50 X 1 /1000= 0.05 Q = ΔH X no of mol = 57.3 X 0.05 = 2.865 kJ // 2865 J	1 1 1												
	(ii)	2865 = 100 X 4.2 X θ θ = 2865 ÷ 420 = 6.8 °C (unit must correct)	1 1 1												
	(iii)	Some of heat is lost to the surrounding Heat is absorbed by polystyrene cup	1 1												
	(c)	<table><tr><th>A</th><th>B</th></tr><tr><td>The reaction is exothermic// Heat releases to the surrounding during the reaction</td><td>The reaction is endothermic// Heat absorbed from the surrounding during the reaction</td></tr><tr><td>During the reaction, heat releases to the surrounding is x kJ per mol</td><td>During the reaction, heat absorbed from the surrounding is y kJ per mol</td></tr><tr><td>The energy content in reactant is higher than energy content in product</td><td>The energy content in reactant is lower than energy content in product</td></tr><tr><td>The temperature increases during the reaction</td><td>The temperature decreases during the reaction</td></tr><tr><td>Heat released during the formation of bond in product is higher than heat absorbed during the breaking of bond in reactant</td><td>Heat absorbed during the breaking of bond in reactant is higher than heat released during the formation of bond in product</td></tr></table>	A	B	The reaction is exothermic// Heat releases to the surrounding during the reaction	The reaction is endothermic// Heat absorbed from the surrounding during the reaction	During the reaction, heat releases to the surrounding is x kJ per mol	During the reaction, heat absorbed from the surrounding is y kJ per mol	The energy content in reactant is higher than energy content in product	The energy content in reactant is lower than energy content in product	The temperature increases during the reaction	The temperature decreases during the reaction	Heat released during the formation of bond in product is higher than heat absorbed during the breaking of bond in reactant	Heat absorbed during the breaking of bond in reactant is higher than heat released during the formation of bond in product	1 1 1 1 1
A	B														
The reaction is exothermic// Heat releases to the surrounding during the reaction	The reaction is endothermic// Heat absorbed from the surrounding during the reaction														
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The temperature increases during the reaction	The temperature decreases during the reaction														
Heat released during the formation of bond in product is higher than heat absorbed during the breaking of bond in reactant	Heat absorbed during the breaking of bond in reactant is higher than heat released during the formation of bond in product														
		TOTAL	20												

SECTION C

SULIT

4541/2

No		Essay Section C	Mark							
9	a(i)	$S + O_2 \longrightarrow SO_2$	2							
				Subtotal 2						
	(ii)	Neutralised sulphur dioxide gas released with calcium oxide/calcium hydroxide/slaked lime/ quick lime	1 1							
				Subtotal 2						
		<table><tr><td>X: ceramics</td><td>High melting point/high heat resistant</td></tr><tr><td>Y: Bronze</td><td>Does not corrode easily/shiny</td></tr><tr><td>Z : Fibre glass</td><td>Light and strong</td></tr></table>	X: ceramics	High melting point/high heat resistant	Y: Bronze	Does not corrode easily/shiny	Z : Fibre glass	Light and strong		
X: ceramics	High melting point/high heat resistant									
Y: Bronze	Does not corrode easily/shiny									
Z : Fibre glass	Light and strong									
				Subtotal 6						
	b	Alloy is Bronze its pure metal is copper or Alloy is Brass its pure metal is copper Materials: steel ball, bronze block/brass block, copper block Apparatus: 1 kg weight, ruler, retort stand with ruler Procedure: 1. A steel ball bearing is taped on to the surface of the copper block 2. A 1 kg weight is hung at the retort stand 50cm above 3. The weight is dropped onto the steel ball 4. The diameter of the dent is measured 5. Step 1-4 are repeated using the bronze block/brass block Observation: Diameter of the for bronze/brass is smaller than copper (vice versa) Conclusion: Bronze/brass is harder than copper	1 1 1 5 1 1							
				Subtotal 10						
TOTAL				20						

Pertis

8

SPM Kimia Paper 12) Jawapan 2012

No	Description	Mark															
10	<p>(a)</p> <p>(i) Chemical equation: $\text{H}_2\text{SO}_4 + \text{CuO} \rightarrow \text{CuSO}_4 + \text{H}_2\text{O}$</p> <p>(ii) No of Mol of H_2SO_4 : $0.2 \times 50 / 1000 = 0.01$</p> <p>From the equation: 1 mol H_2SO_4 produce 1 mol CuSO_4 0.01 mol H_2SO_4 produce 0.01 mol CuSO_4</p> <p>Mass of CuSO_4 produced: $0.01 \times 160 = 1.6 \text{ g}$</p>	<p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>4</p>															
	<p>(b)</p> <table border="1"> <thead> <tr> <th>Aspect</th><th>NaCl</th><th>AgCl</th></tr> </thead> <tbody> <tr> <td>Solubility in water</td><td>Soluble</td><td>Insoluble</td></tr> <tr> <td>Name of reaction</td><td>Neutralisation reaction</td><td>Precipitation / Double decomposition reaction</td></tr> <tr> <td>Reactants</td><td>HCl and NaOH</td><td>AgNO_3 and NaCl/ KCl/ NH_4Cl/ HCl</td></tr> <tr> <td>Chemical equation</td><td>$\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$</td><td>$\text{AgNO}_3 + \text{KCl} \rightarrow \text{AgCl} + \text{KNO}_3$</td></tr> </tbody> </table>	Aspect	NaCl	AgCl	Solubility in water	Soluble	Insoluble	Name of reaction	Neutralisation reaction	Precipitation / Double decomposition reaction	Reactants	HCl and NaOH	AgNO_3 and NaCl/ KCl/ NH_4Cl / HCl	Chemical equation	$\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$	$\text{AgNO}_3 + \text{KCl} \rightarrow \text{AgCl} + \text{KNO}_3$	<p>1</p> <p>1</p> <p>(1+1)</p> <p>(1+1)</p> <p>6</p>
Aspect	NaCl	AgCl															
Solubility in water	Soluble	Insoluble															
Name of reaction	Neutralisation reaction	Precipitation / Double decomposition reaction															
Reactants	HCl and NaOH	AgNO_3 and NaCl/ KCl/ NH_4Cl / HCl															
Chemical equation	$\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$	$\text{AgNO}_3 + \text{KCl} \rightarrow \text{AgCl} + \text{KNO}_3$															
	<p>(c)</p> <p>Reactants: (0.1 – 2.0 mol dm^{-3}) barium nitrate solution, $\text{Ba}(\text{NO}_3)_2$, (0.1 – 2.0 mol dm^{-3}) sodium sulphate, Na_2SO_4 solution or any soluble sulphates or (0.1 – 2.0 mol dm^{-3}) sulphuric acid</p> <p>Procedure: 1. (20 - 100 cm^3) of barium nitrate, $\text{Ba}(\text{NO}_3)_2$ is added to (20 - 100 cm^3) of sodium sulphate, Na_2SO_4. 2. Stir the solution mixture. 3. A white precipitate barium sulphate, BaSO_4, is formed. 4. Filter the solution mixture. 5. Wash / rinse the residue/solid/salt with distilled water. 6. Press the residue with a few pieces of filter papers to dry them.</p> <p>Chemical equation: $\text{Ba}(\text{NO}_3)_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{NaNO}_3$ or $\text{Ba}^{2+} + \text{SO}_4^{2-} \rightarrow \text{BaSO}_4$</p>	<p>(1+1)</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>(1+1)</p> <p>10</p>															
Total		20															

SULIT

4541/3

4541/3
Chemistry
Kertas 3
2012
1 ½ jam

NAMA:

TINGKATAN:



**PEPERIKSAAN PERCUBAAN BERSAMA
SIJIL PELAJARAN MALAYSIA 2012**

**ANJURAN
PERSIDANGAN KEBANGSAAN PENGETUA-PENGETUA
SEKOLAH MENENGAH CAWANGAN PERLIS**

CHEMISTRY

Kertas 3

Satu jam tiga puluh minit

JANGAN BUKA KERTAS SOALAN INI SEHINGGA DIBERITAHU

1. *Tuliskan nama dan tingkatan anda pada ruang yang disediakan.*
2. *Kertas soalan ini adalah dalam dwibahasa.*
3. *Soalan dalam bahasa Inggeris mendahului soalan yang sepadan dalam bahasa Melayu.*
4. *Calon dibenarkan menjawab keseluruhan atau sebahagian soalan sama ada dalam bahasa Melayu atau bahasa Inggeris.*
5. *Calon dikehendaki membaca maklumat di halaman 10.*

Kod Pemeriksa		
Soalan	Markah Penuh	Markah diperolehi
1	18	
2	15	
3	17	
JUMLAH		

Kertas soalan ini mengandungi 10 halaman bercetak

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2

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Answer all the questions.
Jawab semua soalan.

1. Diagram 1.1 shows the apparatus set-up for an experiment to construct the electrochemical series based on the potential differences between two different metals.

Rajah 1.1 menunjukkan susunan radas dalam satu eksperimen untuk membina siri elektrokimia berdasarkan beza keupayaan antara dua logam yang berlainan.

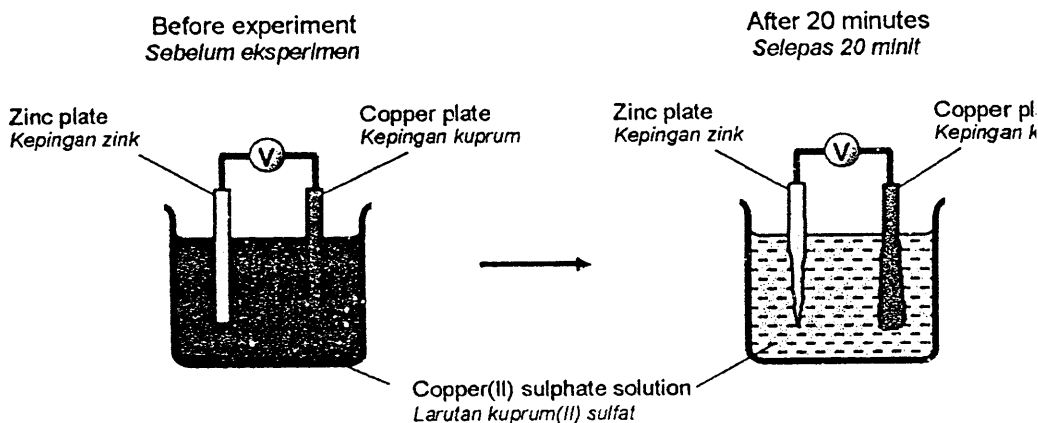


Diagram 1.1
Rajah 1.1

1(a)

- (a) Based on Diagram 1.1, state three observations.
Berdasarkan Rajah 1.1, nyatakan tiga pemerhatian.

.....

.....

.....

[3 marks]

- (b) Experiment is repeated by replacing zinc plate with metals P, Q and R while keeping the copper plate as the positive terminal. Fresh copper(II) sulphate solution is used in each experiment.

Eksperimen diulangi dengan menggantikan kepingan zink dengan logam P, Q dan R sementara mengekalkan kepingan kuprum sebagai terminal positif. Larutan kuprum(II) sulfat yang baru digunakan bagi setiap eksperimen.

[Lihat halaman seterusnya]

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Diagram 1.2 shows the potential differences of different pairs of metals in this experiment.

Rajah 1.2 menunjukkan beza keupayaan bagi pasangan logam yang berlainan dalam eksperimen ini.

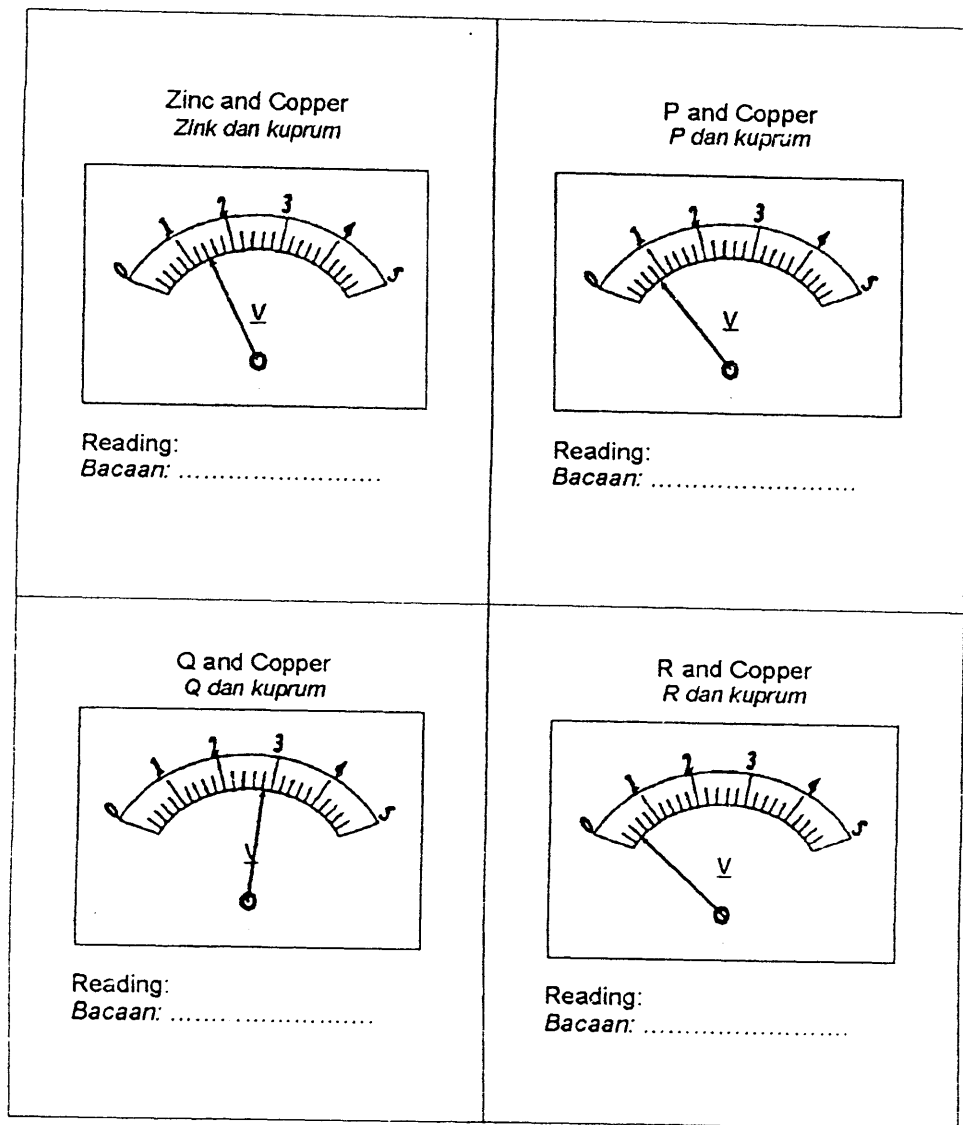


Diagram 1.2
 Rajah 1.2

Record the voltmeter readings in the spaces provided in Diagram 1.2.
 Catatkan bacaan voltmeter pada ruang yang disediakan dalam Rajah 1.2.

[3 marks]

1(b)



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- (c) Construct a table to record the voltmeter readings of four pairs of metals.
Bina satu jadual untuk merekodkan bacaan voltmeter untuk empat pasangan logam.

1(c)

--

[3 marks]

1(d)

--

- (d) Based on the table constructed in 1(c), arrange the metals P, Q, R, zinc and copper in ascending order of the electropositivity of metals.

Berdasarkan jadual yang dibina dalam 1 (c), susunkan logam P, Q, R, zink dan kuprum mengikut keelektropositifan logam secara tertib menaik.

.....
[3 marks]

1(e)

--

- (e) Predict the voltage produced if pair of metals P and Q is placed in copper(II) sulphate solution.

Ramalkan nilai voltan yang terhasil sekiranya pasangan logam P and logam Q diletakkan di dalam larutan kuprum(II) sulfat.

.....
[3 marks]

- (f) Diagram 1.3 shows the change in the mass of zinc plate after 20 minutes.
 Rajah 1.3 menunjukkan perubahan jisim kepingan zink selepas 20 minit.

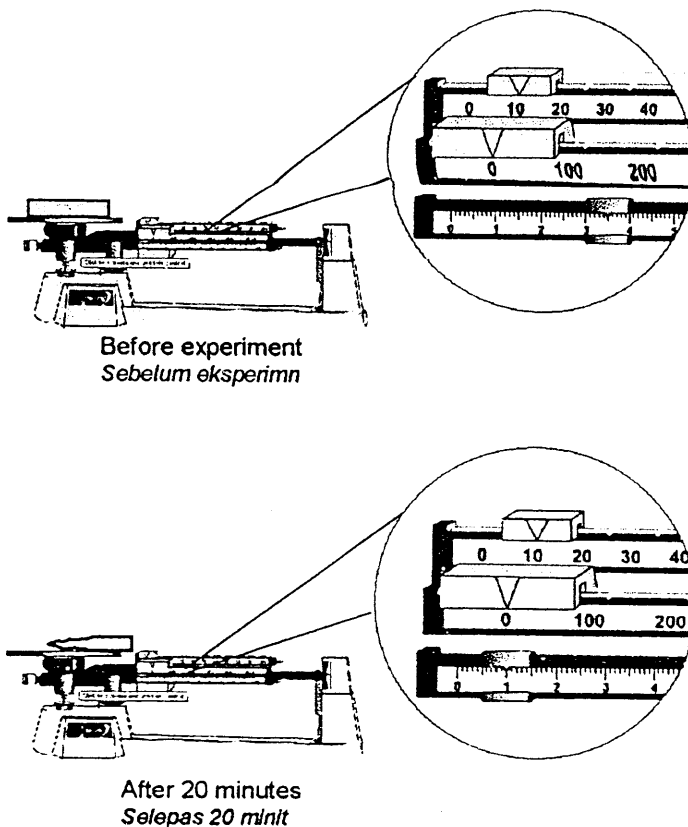
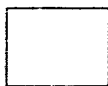


Diagram 1.3
 Rajah 1.3

Based on Diagram 1.3, state the relationship between time taken and the mass of zinc plate remains.

Berdasarkan Rajah 1.3, nyatakan hubungan antara masa yang diambil dengan jisim kepingan zink yang tertinggal.

1(f)



.....

.....

.....

[3 marks]

SULIT

6

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2. Diagram 2 shows the set-up apparatus of the experiment to investigate the effect of other metal on the rusting of iron.
Rajah 2 menunjukkan susunan radas bagi eksperimen untuk mengkaji kesan logam lain ke atas pengurangan besi.

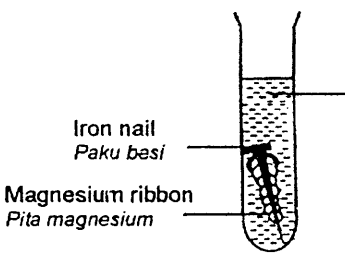
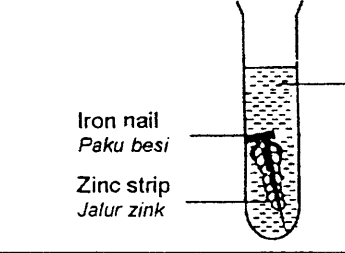
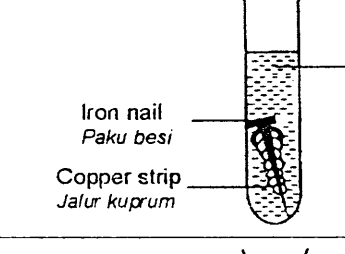
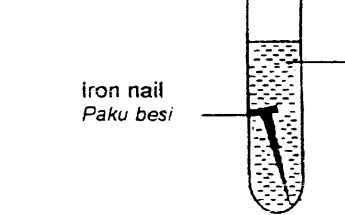
Test tube	Set-up of apparatus <i>Susunan radas</i>	Observation <i>Pemerhatian</i>
I	 <p>Jelly solution + potassium hexacyanoferrate(III) solution + phenolphthalein <i>Larutan agar-agar + larutan kalium heksasianoferat (III) dan fenolftalein</i></p> <p>Iron nail <i>Paku besi</i></p> <p>Magnesium ribbon <i>Pita magnesium</i></p>	<p>Pink colour is produced <i>Warna merah jambu terbentuk</i></p>
II	 <p>Jelly solution + potassium hexacyanoferrate(III) solution + phenolphthalein <i>Larutan agar-agar + larutan kalium heksasianoferat (III) dan fenolftalein</i></p> <p>Iron nail <i>Paku besi</i></p> <p>Zinc strip <i>Jalur zink</i></p>	<p>Pink colour is produced <i>Warna merah jambu terbentuk</i></p>
III	 <p>Jelly solution + potassium hexacyanoferrate(III) solution + phenolphthalein <i>Larutan agar-agar + larutan kalium heksasianoferat (III) dan fenolftalein</i></p> <p>Iron nail <i>Paku besi</i></p> <p>Copper strip <i>Jalur kuprum</i></p>	<p>Blue colour is produced <i>Warna biru terbentuk</i></p>
IV	 <p>Jelly solution + potassium hexacyanoferrate(III) solution + phenolphthalein <i>Larutan agar-agar + larutan kalium heksasianoferat (III) dan fenolftalein</i></p> <p>Iron nail <i>Paku besi</i></p>	<p>Blue colour is produced <i>Warna biru terbentuk</i></p>

Diagram 2
Rajah 2

For
Examiner's
use

SULIT

7

4541/3

2(a)

- (a) State the inference for each of the following test tubes.
Nyatakan inferens bagi setiap tabung uji berikut.

(i) Test tube I

Tabung uji I :

(ii) Test tube II

Tabung uji II :

(iii) Test tube III

Tabung uji III :

[3 marks]

- (b) For this experiment, state :
Bagi eksperimen ini, nyatakan :

(i) The manipulated variable
Pembolehubah dimanipulasikan

.....

(ii) The responding variable
Pembolehubah bergerak balas

.....

2(b)

(iii) The fixed variable
Pembolehubah dimalarkan

.....

[3 marks]

2(c)

- (c) State the hypothesis for the experiment.
Nyatakan hipotesis bagi eksperimen ini.

.....

.....

[3 marks]

2(d)

- (d) State the operational definition for rusting of iron.
Nyatakan definisi secara operasi bagi pengurangan besi.

.....

[3 marks]

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Examiner's use
SULIT

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- (e) Based on this experiment, classify the metals which in contact with the iron nail into metal that is more electropositive than iron and metal that is less electropositive than iron.

Berdasarkan eksperimen ini, kelaskan logam yang bersentuhan dengan paku besi kepada logam yang lebih elektropositif daripada besi dan logam yang kurang elektropositif daripada besi.

Metal that is more electropositive than Iron <i>Logam yang lebih elektropositif daripada besi</i>	Metal that is less electropositive than Iron <i>Logam yang kurang elektropositif daripada besi</i>

2(e)



[3 marks]

3.

Ionic compound can conduct electricity in aqueous and molten states but covalent compound cannot conduct electricity in all states.

Sebatian ion boleh mengkonduksikan elektrik dalam keadaan leburan dan larutan akueus tetapi sebatian kovalen tidak boleh mengkonduksikan elektrik dalam semua keadaan.

Using lead(II) bromide as an example of ionic compound and naphthalene as an example of covalent compound, plan a laboratory experiment to compare the electrical conductivity of molten ionic and covalent compounds.

Dengan menggunakan plumbum(II) bromida sebagai contoh sebatian ion and naftalena sebagai contoh sebatian kovalen, rancangkan satu eksperimen untuk membandingkan kekonduksian elektrik leburan sebatian ion dan sebatian kovalen.

Your planning should include the following aspects:

Perancangan anda hendaklah mengandungi aspek berikut:

- (a) Problem statement
Pernyataan masalah
- (b) All the variables
Semua pemboleh ubah
- (c) Hypothesis
Hipotesis
- (d) List of materials and apparatus
Senarai bahan dan radas
- (e) Procedure
Prosedur
- (f) Tabulation of data
Penjadualan data

[17 marks]

END OF QUESTION PAPER
KERTAS SOALAN TAMAT

SULIT

Pertis

1
Kimia Paper (3) Jawapan

MARKING SCHEME FOR PAPER 3 2012 Trial Pertis

Question	Rubric	Score
1(a)	<p><i>Able to state all three observations correctly</i></p> <p><u>Sample answers:</u></p> <ol style="list-style-type: none"> 1. Zinc plate becomes thinner // 2. Brown solid is deposited // 3. Blue solution changes paler // 4. Pointer of voltmeter deflects 	3
	<p><i>Able to state any 2 observations correctly or three observations less correctly</i></p> <p><u>Sample answers:</u></p> <ol style="list-style-type: none"> 1. Zinc plate dissolves / corrodes // 2. Copper plate becomes thicker // 3. Copper(II) sulphate solution changes paler // Solution changes paler // 	2
	<p><i>Able to state any 1 answer correctly or give the idea of the observations</i></p> <p><u>Sample answers:</u></p> <ol style="list-style-type: none"> 1. some of the zinc disappear // 2. brown precipitate is produced // 3. Solution turns colourless // 4. voltmeter deflects 	1
	<i>No response or wrong response</i>	0

Question	Rubric	Score
1(b)	<p><i>Able to state all the voltmeter readings accurately with correct unit</i></p> <p><u>Sample answer:</u></p> <p>Zinc & copper : 1.4 V P & copper : 0.8 V Q & copper : 2.8 V R & copper : 0.4 V</p>	3
	<p><i>Able to state all the voltmeter readings less accurately /without unit</i></p> <p><u>Sample answer:</u></p> <p>Zinc & copper : 1.40 V // 1.4 P & copper : 0.80 V // 0.8 Q & copper : 2.80 V // 2.8 R & copper : 0.40 V // 0.4</p>	2
	<p><i>Able to state the voltmeter readings</i></p> <p><u>Sample answer:</u></p> <p>Zinc & copper : 1.40 // 1.2 P & copper : 0.80 // 0.4 Q & copper : 2.80 // 2.4 R & copper : 0.40 // 0.2</p>	1
	<i>No response or wrong response</i>	0

SULIT

2

Question	Rubric	Score																				
1(c)	<p><i>Able to construct a table to record the voltmeter readings with one decimal place and with correct unit</i></p> <p><u>Sample answer:</u></p> <table><tr><th>Pair of metals</th><th>Voltmeter reading (V)</th></tr><tr><td>Zinc & copper</td><td>1.4</td></tr><tr><td>P & copper</td><td>0.8</td></tr><tr><td>Q & copper</td><td>2.8</td></tr><tr><td>R & copper</td><td>0.4</td></tr></table>	Pair of metals	Voltmeter reading (V)	Zinc & copper	1.4	P & copper	0.8	Q & copper	2.8	R & copper	0.4	3										
Pair of metals	Voltmeter reading (V)																					
Zinc & copper	1.4																					
P & copper	0.8																					
Q & copper	2.8																					
R & copper	0.4																					
	<p><i>Able to construct a table to record the voltmeter readings without unit</i></p> <p><u>Sample answer:</u></p> <table><tr><th>Pair of metals</th><th>Voltmeter reading</th></tr><tr><td>Zinc & copper</td><td>1.4</td></tr><tr><td>P & copper</td><td>0.8</td></tr><tr><td>Q & copper</td><td>2.8</td></tr><tr><td>R & copper</td><td>0.4</td></tr></table> <p style="text-align: center;"><i>or</i></p> <table><tr><th>Pair of metals</th><th>Voltmeter reading</th></tr><tr><td>Zinc & copper</td><td>1.4 V</td></tr><tr><td>P & copper</td><td>0.8 V</td></tr><tr><td>Q & copper</td><td>2.8 V</td></tr><tr><td>R & copper</td><td>0.4 V</td></tr></table>	Pair of metals	Voltmeter reading	Zinc & copper	1.4	P & copper	0.8	Q & copper	2.8	R & copper	0.4	Pair of metals	Voltmeter reading	Zinc & copper	1.4 V	P & copper	0.8 V	Q & copper	2.8 V	R & copper	0.4 V	2
Pair of metals	Voltmeter reading																					
Zinc & copper	1.4																					
P & copper	0.8																					
Q & copper	2.8																					
R & copper	0.4																					
Pair of metals	Voltmeter reading																					
Zinc & copper	1.4 V																					
P & copper	0.8 V																					
Q & copper	2.8 V																					
R & copper	0.4 V																					
	<p><i>Able to construct a table to record the voltmeter readings with either one correct heading</i></p>	1																				
	<p><i>No response or wrong response</i></p>	0																				

Question	Rubric	Score
1(d)	<p><i>Able to arrange the metals P, Q, R, zinc and copper in ascending order of the electropositivity of metals</i></p> <p><u>Sample answers:</u> Copper, R, P, Zinc, Q</p>	3
	<p><i>Able to arrange all metals in descending order of the electropositivity of metals // arrange at least 3 metals in correct sequence according to ascending order of the electropositivity of metals.</i></p> <p><u>Sample answers:</u> Q, Zinc, P, R, Copper // R, P, zinc, Q, copper</p>	2
	<p><i>Able to arrange at least 2 metals in correct sequence either in ascending order or in descending order of the electropositivity of metals</i></p> <p><u>Sample answers:</u> Copper, R, Q zinc, P // Q, zinc, R, copper, P</p>	1
	<i>No response or wrong response</i>	0

SULIT Perlis

3

Kimia Paper (3) Jawapan 2012

Question	Rubric	Score
1(e)	<i>Able to predict the voltage produced accurately with correct unit</i> <u>Sample answer:</u> 2.0 V	3
	<i>Able to predict the voltage produced accurately without correct unit</i> <u>Sample answer:</u> 2.0	2
	<i>Able to predict the voltage produced less accurately</i> <u>Sample answer:</u> Less than 2.8 V // More than 0.8 V	1
	<i>No response or wrong response</i>	0

Question	Rubric	Score
1(f)	<i>Able to state the relationship between the time taken for the reaction to occur and the mass of zinc plate correctly or vice-versa</i> <u>Sample answer:</u> The longer the time taken for the reaction to occur, the lesser the mass of zinc plate remains // the shorter the time taken for the reaction to occur, the more the mass of zinc plate remains	3
	<i>Able to state the relationship between the time taken for the reaction to occur and the mass of zinc plate less correctly</i> <u>Sample answer:</u> The higher the time taken for the reaction to occur, the lesser the mass of zinc plate remains // When the mass of zinc plate decreases, the time taken for the reaction to occur increases	2
	<i>Able to give an idea of the relationship between the time taken for the reaction to occur and the mass of zinc plate</i> <u>Sample answer:</u> The mass of zinc plate is affected by the time taken	1
	<i>No response or wrong response</i>	0

SULIT

4

Question	Rubric	Score
2(a)	<p><i>Able to state all the inferences correctly</i></p> <p><u>Sample answers:</u> Test tube I : iron nail does not rust Test tube II : iron nail does not rust Test tube III : iron nail rust</p>	3
	<p><i>Able to state all the inferences less correctly</i></p> <p><u>Sample answers:</u> Test tube I : OH^- ion is present Test tube II : OH^- ion is present Test tube III : Fe^{2+} ion is present</p>	2
	<p><i>Able to give an idea for the inferences</i></p> <p><u>Sample answers:</u> Test tube I : iron nail rust // Fe^{2+} ion is present Test tube II : iron nail rust // Fe^{2+} ion is present Test tube III : iron nail does not rust // OH^- ion is present</p>	1
	<i>No response or wrong response</i>	0

Question	Rubric	Score
2(b)	<p><i>Able to state all the three variables correctly</i></p> <p><u>Sample answer:</u> Manipulated variable: Type of metals in contact with iron nail</p> <p>Responding variable: Rusting of iron nail // colour produced</p> <p>Constant variable: Type of nail // iron nail</p>	3
	<p><i>Able to state three variables less correctly or any two variables correctly</i></p> <p><u>Sample answer:</u> Manipulated variable: Type of metals // pair of metals</p> <p>Responding variable: Rate of rusting</p> <p>Constant variable: Type of iron nail</p>	2
	<i>Able to state any one variable correctly</i>	1
	<i>No response or wrong response</i>	0

SULIT Perlis

5

Kimia Paper (3) Jawapan 2012

Question	Rubric	Score
2(c)	<p><i>Able to state the relationship between the manipulated variable and the responding variable with correct direction.</i></p> <p><u>Sample answer:</u> When iron nail is in contact with a more electropositive metal, rusting does not occur; when iron nail is in contact with a less electropositive metal, rusting occurs</p>	3
	<p><i>Able to state the relationship between the manipulated variable and responding variable but less accurate in stating the direction.</i></p> <p><u>Sample answer:</u> When iron nail is in contact with a more electropositive metal, rusting does not occur // When iron nail is in contact with a less electropositive metal, rusting occurs //</p> <p>Rusting occurs when iron nail is in contact with a less electropositive metal but rusting does not occur when iron nail is in contact with a more electropositive metal //</p> <p>When iron nail is in contact with a more electropositive metal, rusting is prevented; when iron nail is in contact with a less electropositive metal, rusting is speeded up//</p> <p>Iron nail rusts when copper is used but iron nail does not rust with zinc and magnesium are used</p>	2
	<p><i>Able to give an idea of hypothesis</i></p> <p><u>Sample answer:</u> Type of metals coiled on iron nail affects the rusting or iron.</p>	1
	<i>No response or wrong response</i>	0

Question	Rubric	Score
2(d)	<p><i>Able to fulfill the following aspects</i> (i) what have been done (ii) what is observed</p> <p><u>Sample answer:</u> Formation of blue colour when iron nail is coiled with a less electropositive metal</p>	3
	<p><i>Able to state the operational definition less correctly</i></p> <p><u>Sample answer:</u> Formation of rust when iron nail is coiled with a less electropositive metal</p>	2
	<p><i>Able to state an idea for the rusting of iron</i></p> <p><u>Sample answer:</u> Formation of rust / Fe^{2+} ion // oxidation / corrosion of iron</p>	1
	<i>No response or wrong response</i>	0

SULIT

6

Question	Rubric	Score				
2(e)	<p><i>Able to classify all the metals correctly</i></p> <p><u>Sample answer:</u></p> <table><tr><td>Metal that is more electropositive than iron</td><td>Metal that is less electropositive than iron</td></tr><tr><td>zinc magnesium</td><td>copper</td></tr></table>	Metal that is more electropositive than iron	Metal that is less electropositive than iron	zinc magnesium	copper	3
Metal that is more electropositive than iron	Metal that is less electropositive than iron					
zinc magnesium	copper					
	<i>Able to classify two metals correctly in both categories or in vice-versa</i>	2				
	<i>Able to classify one metal correctly</i>	1				
	<i>No response</i>	0				

SULIT Perlis

7

Kimia Paper (3) Jawapan 2012

Question	Rubric	Score
3(a)	Able to state the problem statement of the experiment correctly <u>Sample answer</u> How does the type of compounds affect the electrical conductivity? // Does lead(II) bromide conduct electricity in molten state but Naphthalene does not conduct electricity in molten state ?	3
	Able to state the problem statement of the experiment less correctly <u>Sample answer</u> Does the lead(II) bromide conduct electricity in molten state ? // Does the naphthalene conduct electricity in molten state ?	2
	Able to give an idea of the problem statement <u>Sample answer</u> To investigate / compare the electrical conductivity of ionic and covalent compound	1
	No response or wrong response	0

Question	Rubric	Score
3(b)	Able to state the three variables correctly. <u>Sample answer</u> Manipulated variable : lead(II) bromide and naphthalene// Ionic and covalent compounds Responding variable : light up of bulb // deflection of ammeter / voltmeter' needle // electrical conductivity Fixed variable : carbon electrodes	3
	Able to state three variables less correctly or any two variables correctly. <u>Sample answer</u> Manipulated variable : types of substances Responding variable : ammeter / voltmeter reading Fixed variable : mass of substances / bulb / ammeter	2
	Able to state two variables less correctly or any one variable correctly.	1
	No response or wrong response	0

SULIT

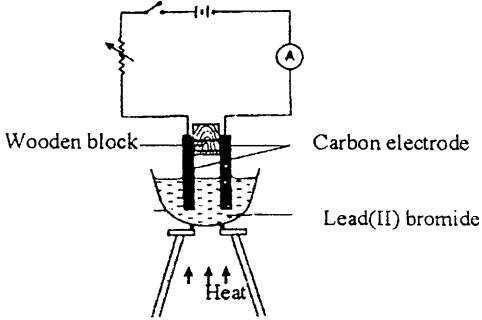
8

Question	Rubric	Score
3(c)	<p><i>Able to suggest a relationship correctly between the MV and the RV with direction</i></p> <p><u>Sample answer</u> Molten lead(II) bromide causes the bulb to light up whereas molten naphthalene does not cause the bulb to light up // Molten lead(II) bromide can conduct electricity whereas molten naphthalene cannot conduct electricity</p>	3
	<p><i>Able to suggest a relationship between the MV and the RV</i></p> <p><u>Sample answer</u> Molten lead(II) bromide // molten ionic compound causes the bulb to light up / conducts electricity // Molten naphthalene // molten covalent compound does not cause the bulb to light up / conducts electricity // Lead(II) bromide can conduct electricity whereas naphthalene cannot conduct electricity //</p>	2
	<p><i>Able to suggest an idea of hypothesis</i></p> <p><u>Sample answer</u> Lead(II) bromide / naphthalene affects the electrical conductivity</p>	1
	No response or wrong response	0

Question	Rubric	Score
3(d)	<p><i>Able to give complete list of substances and apparatus</i></p> <p><u>Sample answer</u> 2 substances : Lead(II) bromide , naphthalene 8 apparatus : batteries , carbon electrodes , bulb/ammeter , crucible , Bunsen burner , tripod stand , pipe-clay triangle , connecting wires</p> <p>[can refer to labelled diagram or procedure but only 1 substance and 2 apparatus]</p>	3
	<p><i>Able to give an incomplete list of substances and apparatus</i></p> <p><u>Sample answer</u> 2 substances : Lead(II) bromide , naphthalene 8 apparatus : batteries , carbon electrodes , bulb/ammeter , [any suitable container], Bunsen burner , connecting wires</p> <p>[can refer to labelled diagram or procedure but only 1 substance and 1 apparatus]</p>	2
	<p><i>Able to give at least one substance and batteries , carbon electrodes , bulb/ammeter, connecting wires</i></p>	1
	No response or wrong response	0

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Question	Rubric	Score
3(e)	<p><i>Able to list all the steps correctly</i></p> <p><u>Sample answer</u></p> <ol style="list-style-type: none"> 1. A crucible is filled with lead(II) bromide solid until it is half full. 2. Two carbon electrodes are immersed into lead(II) bromide and carbon electrodes are connected to batteries and bulb. 3. Record observation 4. The lead(II) bromide solid is heated until its melts. 5. Record observation 6. Repeat steps 1 to 5 using naphthalene. <p>[If description in procedure incomplete, can refer to diagram]</p> 	3
	<i>Able to list steps 1, 2, 4 and 6</i>	2
	<i>Able to give steps 2 and 4</i>	1
	No response or wrong response	0

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10

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Question	Rubric	Score						
3(f)	<p><i>Able to fulfill the following aspects:</i></p> <ol style="list-style-type: none"><i>correct headings</i><i>Complete list of manipulated variables</i> <p><u>Sample answer</u></p> <table><tr><th>Type of compound</th><th>Observation</th></tr><tr><td>Lead(II) bromide</td><td></td></tr><tr><td>naphthalene</td><td></td></tr></table>	Type of compound	Observation	Lead(II) bromide		naphthalene		2
Type of compound	Observation							
Lead(II) bromide								
naphthalene								
	<p><i>Able to tabulate the data with the following aspects:</i></p> <ol style="list-style-type: none"><i>one of the suitable headings</i><i>one name of the compound</i> <p><u>Sample answer</u></p> <table><tr><th>Substance</th><th>Observation</th></tr><tr><td>Lead(II) bromide</td><td></td></tr></table>	Substance	Observation	Lead(II) bromide		1		
Substance	Observation							
Lead(II) bromide								
	No response or wrong response or empty table	0						

END OF MARKING SCHEME